

Cambridge IGCSE™

CO-ORDINATED SCIENCES**0654/52**

Paper 5 Practical Test

February/March 2025

MARK SCHEME

Maximum Mark: 60

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the February/March 2025 series for most Cambridge IGCSE, Cambridge International A and AS Level components, and some Cambridge O Level components.

This document consists of **10** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptions for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.

2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.

3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).

4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.

5 'List rule' guidance

For questions that require *n* responses (e.g. State **two** reasons ...):

- The response should be read as continuous prose, even when numbered answer spaces are provided.
- Any response marked *ignore* in the mark scheme should not count towards *n*.
- Incorrect responses should not be awarded credit but will still count towards *n*.
- Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
- Non-contradictory responses after the first *n* responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

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Question	Answer	Marks
1(a)	full set of results for boiling tube ; full set of results for test-tube ; greater temperature fall in test-tube ; all values to nearest 0.5 ; temperature reduces over time in both tubes ;	5
1(b)(i)	axes right way round and labelled with quantity and unit ; e.g.: y-axis temperature/T and °C and x axis time/t and minutes/min(s) sensible linear scale and plots cover at least half the grid ; plots correct for boiling tube $\pm$ half small square ; plots correct for test tube $\pm$ half small square ; curve of best fit for each test-tube and labelled ;	5
1(b)(ii)	correct reading from graph ;	1
1(c)(i)	correct calculations ;	1
1(c)(ii)	larger animals lose heat at a lower rate / more slowly / smaller animals lose heat at a larger rate / more quickly ;	1
1(d)	exclude anomalies/ reduce random error ;	1

Question	Answer	Marks
2(a)(i)	A blue-black ; B brown ;	2
2(a)(ii)	starch present in A and no starch present in B ;	1
2(a)(iii)	light needed for starch production ORA ;	1
2(b)(i)	colour would mask test result / to see colour of iodine solution more clearly ;	1
2(b)(ii)	hot water then hot ethanol then water rinse ;	1

Question	Answer	Marks
3(a)	grey and solid :	1
3(b)(i)	any two from: effervesces ; moves up and down ; forms a colourless solution ;	2
3(b)(ii)	lighted splint ; pops and hydrogen ;	2
3(c)	white ppt and limewater ;	1
3(d)	white ppt and limewater ;	1

Question	Answer	Marks
3(e)	white ppt and insoluble in excess ;	1
3(f)	calcium ion/ Ca^{2+} and J is calcium/Ca ;	1
3(g)	nitric acid and silver nitrate ; white ppt and chloride ;	2
3(h)	blue flame is hotter ; doesn't mask colour of metal ion in flame ;	2

Question	Answer	Marks
4	apparatus and safety syringe and volume / amount gas measured in the method ; balance and mass / amount solid measured in the method ; method of heating and airtight apparatus ; goggles to protect eyes from sodium hydroxide / gloves to protect skin/hands from burning/hot apparatus ; method add ammonium sulfate to aqueous sodium hydroxide and collect gas given off (as the reaction finishes) for at least 2 masses/amounts of ammonium sulfate; measurements and validity volume of ammonia ; mass of ammonium sulfate ; do each mass / amount more than once to identify / exclude anomalies ; five different masses / amounts ; control volume of sodium hydroxide ; concentration of sodium hydroxide ; temperature ; processing for conclusion plot graph of volume against mass ; straight line – linear relationship / straight line through origin – proportional / increasing diagonal line as T increases then volume increases ; when mass increases does volume increase, decrease or stay the same ;	7

Question	Answer	Marks
5(a)(i)	5.5–6.5 ;	1
5(a)(ii)	angle recorded ;	1
5(a)(iii)	other 4 angles present ; All angles increase as height increases ;	2
5(b)	the bigger the height the bigger the angle ;	1
5(c)	video the experiment and play back in slow motion or pause / clamp in position once bottle starts to fall/ do the experiment near a wall and mark position on the wall / place a solid object in front as it starts to fall ;	1
5(d)	value consistent with candidate's table ;	1

Question	Answer	Marks
6(a)(i)	l , w and h measured ; all to nearest 0.1 cm ;	2
6(a)(ii)	volume calculated ; volume to 3 sf ;	2
6(a)(iii)	m recorded to 0.1 g ;	1
6(a)(iv)	density calculated ;	1
6(a)(v)	sides may not be uniform AW ;	1
6(b)(i)	Both volumes recorded both to nearest cm^3 and $V_3 > V_2$;	1
6(b)(ii)	V_c calculated ;	1

Question	Answer	Marks
6(b)(iii)	density calculated ;	1
6(c)(i)	mass of water would not be included in the mass of the modelling clay ;	1
6(c)(ii)	calculation showing 10% and statement it is less or more than 10% and conclusion ;	1
6(c)(iii)	density increases and because of dividing by a smaller V/ density is inversely proportional to volume ;	1